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Kinetic Molecular Theory Pogil Answer

POSTULATES OF THE KINETIC MOLECULAR THEORY Gases consist of tiny particles (atoms or molecules). These particles are so small compared with the distance between them that the volume (size) of the individual particles can be assumed to be negligible (zero).

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The Kinetic-Molecular Theory Explains the Behavior of Gases, Part II According to Graham’s law, the molecules of a gas are in rapid motion and the molecules themselves are small. The average distance between the molecules of a gas is large compared to the size of the molecules.

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Answers Kinetic Molecular Theory Pogil

POGIL: Kinetic Molecular Theory. Modified from Foundations of Chemistryby David HansonPage 1 of 3. POGIL: Kinetic Molecular Theory. Learning Objectives. Identify the basic differences between particle behavior in the solid, liquid, and gaseous phases. Develop an understanding of the postulates of the kinetic molecular theory. Identify how temperature affects molecular motion. Apply the kinetic molecular theory to predict the outcome of everyday situations.

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Review Questions Answer KEY 1. Kinetic molecular theory is the theory that explains the motion of solids, liquids, and gases. 2. KMT explains the differences between properties of solids, liquids, and gases by examining how the particles are moving under similar conditions. 3. a. Solid – a phase of matter with a definite shape and volume.

Review Questions Answer KEY

A. Kinetic Molecular Theory (KMT) = the idea that particles of

matter are always in motion and that this motion has consequences. 1) theory developed in the late 19th century to account for the behavior of the atoms and molecules that make up matter 2) based on the idea that particles in all forms of matter are

I. MOLECULES IN MOTION: A.

POSTULATES OF THE KINETIC MOLECULAR THEORY Gases consist of tiny particles (atoms or molecules). These particles are so small, compared with the distances between them that the volume (size) of the individual particles can be assumed to be negligible (zero). The particles are in constant random motion, colliding with the walls of the container.

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9.5 The Kinetic-Molecular Theory – Chemistry

Kinetic Molecular Theory attempts to explain why gas molecules behave as they do. It is based off four central tenets. 1. Gaseous molecules have no volume 2. Collisions between gaseous molecules are elastic (no energy is gained or lost) 3. Gaseous molecules have no attraction with one another 4.

Kinetic Molecular Theory - AP Chemistry : Gas Laws

Postulates of Kinetic Molecular Theory 1. Gases consist of tiny particles (atoms or molecules). 2. These particles are so small compared with the distances between them that the volume (size) of the individual particles can be assumed to negligible (zero). 3.

Element of the Day S

• The average kinetic energy of a collection of gas particles is directly proportional to the Kelvin temperature of the sample. Critical Thinking Questions 1. We have learned that, for an ideal gas, the pressure is directly proportional to the number of moles, if volume and temperature are constant. Explain this in terms of KMT.

9.1 KMT_POGIL_solutions 17.pdf - Kinetic Molecular Theory ...

Worksheet 13 - Molecular Shapes The shapes of molecules can be predicted from their Lewis structures by using the VSEPR (Valence Shell Electron Pair Repulsion) model, which states that electron pairs around a central atoms will assume a geometry that keeps them as far apart from each other as possible This is illustrated by the drawings below Chem 115 POGIL Worksheet - Week 12 Molecular Shapes Why...